

Michael Offutt

"Solubility Rules"

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Solubility rules they're really great they help predict
precipitates
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Alkali Metal ions are really swell
Assume their compounds are always soluble
Those ions are hardly precipitators
In reactions they're usually spectators
And that positive one ion NH_4
It's compounds are soluble for sure
And that negative one ion NO_3
Nitrates are soluble as can be

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Of course we know one thing is always true
There always exceptions to the rules
Sulfates are soluble it's said
Except with Barium, Strontium and Lead
Chlorides, Bromides and Iodides are soluble
But there's some secrets I'm gonna tell
Working in the lab it's plain to see
They're insoluble with silver lead and mercury

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Many ions make precipitates
Including Phosphates Sulfide and Carbonate
Except when paired with Ammonium or Alkali Metal ions
Hey it's hard to make this simple but I'm trying
So consider all of the possibilities
And use the rules to guess the solubilities
The most important thing that you must keep straight
The insoluble combinations precipitate

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